



# AP Chemistry

## Polyatomic Ions

Negative Charge	Ion name and formula
1-	<b>Acetate ion</b> $\text{C}_2\text{H}_3\text{O}_2^-$ <b>Azide</b> $\text{N}_3^-$ <b>Chlorate ion</b> $\text{ClO}_3^-$ Chlorite ion $\text{ClO}_2^-$ Hypochlorite $\text{ClO}^-$ <b>Cyanide ion</b> $\text{CN}^-$ <b>Hydroxide ion</b> $\text{OH}^-$ <b>Nitrate ion</b> $\text{NO}_3^-$ Nitrite ion $\text{NO}_2^-$ <b>Perchlorate ion</b> $\text{ClO}_4^-$ <b>Permanganate ion</b> $\text{MnO}_4^-$ Hydrogen carbonate ion $\text{HCO}_3^-$ Dihydrogen phosphate ion $\text{H}_2\text{PO}_4^-$ <b>Thiocyanate</b> $\text{SCN}^-$ <b>Bromate</b> $\text{BrO}_3^-$ Hypobromite $\text{BrO}^-$ <b>Iodate</b> $\text{IO}_3^-$ Periodate $\text{IO}_4^-$ Hypoiodite $\text{IO}^-$ Hydrogen sulfate $\text{HSO}_4^-$ Hydrogen sulfite $\text{HSO}_3^-$
2-	<b>Carbonate ion</b> $\text{CO}_3^{2-}$ <b>Chromate ion</b> $\text{CrO}_4^{2-}$ <b>Dichromate ion</b> $\text{Cr}_2\text{O}_7^{2-}$ Hexafluorosilicate $\text{SiF}_6^{2-}$ Hydrogen phosphate ion $\text{HPO}_4^{2-}$ <b>Peroxide ion</b> $\text{O}_2^{2-}$ <b>Sulfate ion</b> $\text{SO}_4^{2-}$ Sulfite ion $\text{SO}_3^{2-}$ <b>Oxalate</b> $\text{C}_2\text{O}_4^{2-}$ <b>Selenate</b> $\text{SeO}_4^{2-}$ <b>Silicate</b> $\text{SiO}_3^{2-}$ <b>Tartrate</b> $\text{C}_4\text{H}_4\text{O}_6^{2-}$ <b>Tetraborate</b> $\text{B}_4\text{O}_7^{2-}$ Thiosulfate $\text{S}_2\text{O}_3^{2-}$
3-	<b>Phosphate ion</b> $\text{PO}_4^{3-}$ Phosphite $\text{PO}_3^{3-}$ <b>Arsenate ion</b> $\text{AsO}_4^{3-}$ <b>Borate</b> $\text{BO}_3^{3-}$

Positive charge	Ion name and formula
1+	<b>Ammonium ion</b> $\text{NH}_4^+$
2+	<b>Mercury(I) ion</b> $\text{Hg}_2^{2+}$

### Common Acid Names

$\text{HC}_2\text{H}_3\text{O}_2$	acetic acid	$\text{HNO}_3$	nitric acid
$\text{CH}_3\text{COOH}$	acetic acid	$\text{H}_3\text{PO}_4$	phosphoric acid
$\text{H}_2\text{CO}_3$	carbonic acid	$\text{H}_2\text{SO}_4$	sulfuric acid
$\text{HCl}$	hydrochloric acid		