

# Chemistry CP

Name: \_\_\_\_\_

Drawing Lewis Structures

Section: \_\_\_\_\_

Use this strategy to write Lewis structures for molecules:

1. Set up the problem
  - Determine total number of valence electrons
  - Arrange atoms (if carbon is present, it will be the central atom)
2. Predict and verify
  - Show pairs of electrons being shared between atoms
  - Put in remaining electrons until each atom has 8 valence electrons (H is content with 2)
  - Replace pairs of shared electrons with dashes
3. Check your work.
  - Make sure the octet rule is satisfied

1. H <sub>2</sub> S	5. SOCl <sub>2</sub> (sulfur is the central atom)
2. SiH <sub>4</sub>	6. CH <sub>4</sub>
3. NH <sub>3</sub>	7. CO
4. PCl <sub>3</sub>	8. CCl <sub>2</sub> F <sub>2</sub>