



6. Find the pH of a 0.09M solution of hydrobromic acid.
  
7. Find the pH and the hydroxide ion concentration of a  $3.5 \times 10^{-3}$ M solution of hydroiodic acid, a strong acid.
  
8. Find the pH and pOH of a  $2.234 \times 10^{-6}$ M solution of hydrochloric acid, a strong acid.
  
9. A solution of chloric acid, a strong acid, has a pOH of 10.32. Determine the hydroxide ion concentration and the pH of the solution.
  
10. A solution of nitric acid, a strong acid has a hydronium ion concentration of  $7.89 \times 10^{-2}$ M. Find the pH and pOH of this solution.

The purpose of this assignment was to:

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Did I:	Circle the appropriate response:		
Explain how I did the problems?	Always	Sometimes	Rarely
Listen while my partner explained?	Always	Sometimes	Rarely
Give my partner positive support?	Always	Sometimes	Rarely
Stay on task during the assignment?	Always	Sometimes	Rarely
Use encouraging and polite words?	Always	Sometimes	Rarely
Record my work on the paper?	Always	Sometimes	Rarely
Demonstrate an understanding of the material?	Yes	No	

Comments: