

Chemistry CP

Name: _____

Pair Tutoring Sheet—Periodicity Vocabulary

Section: _____

Directions:

1. Review the terms and definitions for five minutes.
2. Ask someone to quiz you by giving the definition for each term. You will say the term that matches that definition.
3. If your response is correct, put a check in the box next to that term. If your answer is incorrect, put a minus in the box, and the person quizzing you should explain why the answer is incorrect.
4. Repeat steps 1-3 with a new partner. This time, have your partner give you the term, and you give the definition. Repeat this process until three checks in a row are completed for each item.
5. Have each partner sign his or her name, with a compliment.

Terms	Responses				Definitions
Alkali metals					The elements in group 1
Group					Elements located in the same vertical column in the periodic table; assigned numbers from 1 to 18
Lanthanides					The set of elements from lanthanum through ytterbium
Actinides					The set of elements from actinium through nobelium
Metal					Elements that are hard, shiny, and good conductors
Metalloid					Elements with both metallic and nonmetallic properties
Nonmetal					When forming compounds, frequently gains electrons and can share its outer electrons with other atoms
Octet Rule					An atom with eight electrons in its outermost level is unreactive.
Period					The horizontal rows in the periodic table
Periodic Law					The properties of elements are a periodic function of their atomic numbers
Transition Element					The elements in groups 3 through 12
Triad					Term coined by Dobereiner to describe groups of three related elements
Alkali earth metals					The elements in group 2
Halogens					The elements in group 17
Noble gases					The elements in group 18
Atomic radius					The size of an atom
Valence electrons					The electrons in the outermost s and p orbitals
First ionization energy					The energy required to remove the most loosely held electron in an atom
Ionic radius					The size of an ion
Ionization energy					The energy required to remove an electron from an atom
Shielding effect					Outer electrons do not feel the total nuclear charge because it is partially "screened" by the inner electrons
Electronegativity					The tendency of an atom in a compound to attract electrons to itself

Partner: _____ Compliment: _____

Partner: _____ Compliment: _____

Partner: _____ Compliment: _____