



Acid Nomenclature

Bromfield Honors Chemistry

Identifying acids

- Compound with one or more H^+ ions with an anion
 - Ex. HCl
 - Ex. $\text{HC}_2\text{H}_3\text{O}_2$

Binary Acids

- When the anion does NOT contain oxygen
- prefix *hydro* + **root of the anion's name** – *ic* + the word acid
- Examples:
 - HCl *hydro**chloric** acid*
 - HBr *hydro**bromic** acid*

Oxyacids

- When the anion contains oxygen
- The name will depend on the name of the polyatomic anion. DO NOT use the prefix hydro.
 - ATE ↔ IC
 - ITE ↔ OUS

Examples

- H_2SO_4 the anion is **sulfate**, therefore the acid name will end in **ic**
 - **Sulfuric acid**
- H_2SO_3 the anion is **sulfite**, therefore the name of the acid will end in **ous**
 - **Sulfurous acid**