Name:

Section:

Honors Chemistry

Classifying chemical reactions

Directions:
First, classify the following reactions as synthesis, decomposition, single replacement, double displacement, or combustion. Then, identify the reactions as redox or non-redox.

Reaction	Class	Redox? Y or N
1. $3\ Mg + N_2 \rightarrow Mg_3N_2$		
2. $2\ C_2H_6 + 7\ O_2 \rightarrow 4\ CO_2 + 6\ H_2O$		
3. $2\ H_2O \rightarrow 2\ H_2 + O_2$		
4. $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$		
5. $2\ HgO \rightarrow 2\ Hg + O_2$		
6. $2\ KBr + Cl_2 \rightarrow 2\ KCl + Br_2$		
7. $CaO + H_2O \rightarrow Ca(OH)_2$		
8. $AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$		
9. $2\ H_2O_2 \rightarrow 2\ H_2O + O_2$		
10. $Ca(OH)_2 + H_2SO_4 \rightarrow CaSO_4 + 2\ H_2O$		
11. $CH_4 + 2\ O_2 \rightarrow CO_2 + 2\ H_2O$		
12. $2\ Li + Cu(NO_3)_2 \rightarrow 2\ LiNO_3 + Cu$		