



4. How much heat in joules is lost when 16.7 g of  $\text{H}_2\text{O}(\text{g})$  at  $120.^\circ\text{C}$  are converted to liquid at  $63^\circ\text{C}$ ?
5. Calculate the energy required in kJ to bring 308 g of liquid water at  $78.0^\circ\text{C}$  to steam at  $109.0^\circ\text{C}$ .
6. A 92 g sample of ice at  $-16.0^\circ\text{C}$  is heated to water vapor at  $104^\circ\text{C}$ . Calculate the amount of energy (in joules) required.