

NAME:

## HONORS CHEMISTRY

SECTION:

Pair Tutoring: Liquids and Solutions

### Directions:

1. Review the terms and definitions for five minutes.
2. Ask someone to quiz you by giving the definition for each item. You will say the term that matches that definition.
3. If your response is correct, put a check in the box next to that term. If your answer is incorrect, put a minus in the box, and the person quizzing you should explain why the answer is incorrect.
4. Repeat steps 1-3 with a new partner. This time, have your partner give you the word, and you give the definition. Repeat this process until three checks in a row are completed for each item
5. Have each partner sign his or her name, with a compliment.

Terms	Responses				Definitions
Solution					A homogeneous mixture of two or more substances in a single phase
Soluble					Capable of being dissolved
Solute					The substance that gets dissolved in a solution
Solvent					The dissolving medium in a solution
Miscible					Two liquids that dissolve in each other in any proportion
Immiscible					Two liquids that do not dissolve in each other
Colloid					A heterogeneous mixture consisting of particles that are intermediate in size between those in solutions and suspensions
Suspension					Heterogeneous mixture that appears uniform while being stirred, but separates upon standing, because the particles are large
Electrolyte					Dissolves in water to give a solution that conducts electric current, because it dissociates to form ions
Nonelectrolyte					Dissolves in water to give a solution that does not conduct an electric current
Saturated					Holds the maximum amount of dissolved solute
Unsaturated					Contains less than the maximum amount of dissolved solute
Supersaturated					Contains more than the standard amount of solute
Molarity					Moles of solute per liter of solution
Boiling point					The temperature at which the vapor pressure of a liquid equals the atmospheric pressure
Vapor pressure					The pressure exerted by a gas above its liquid
Triple point					The temperature at which the solid, liquid and gas of a substance coexist
Critical point					Above this temperature and pressure, gases and liquids are indistinguishable
Normal boiling point					The temperature at which the vapor pressure of a liquid equals one atmosphere
Freezing point					The temperature at which a liquid becomes a solid
Colligative properties					Properties that depend on the concentration of solute particles but not their identity
Tyndall effect					Scattering of light beams by colloids and suspensions
Molality					Moles of solute per kilogram of solution
Standard solution					A solution whose concentration is accurately known

Partner: \_\_\_\_\_ Compliment: \_\_\_\_\_

Partner: \_\_\_\_\_ Compliment: \_\_\_\_\_

Partner: \_\_\_\_\_ Compliment: \_\_\_\_\_

# Chemistry CP

Name: \_\_\_\_\_

Pair Tutoring: Solutions

Period: \_\_\_\_\_

<b>Observable Behavior</b>	<b>Self observations</b>	<b>Teacher observations</b>
Entered checks or minus on log sheet		
Recorded compliments		
Spoke only to partner		
Explained incorrect answers to partner		
Gave positive support		
Completed entire assignment		
Demonstrated understanding of the material		

Comments: