

AP Chemistry

Polyatomic Ions

Negative Charge	Ion name and formula
1-	Acetate ion $\text{C}_2\text{H}_3\text{O}_2^-$ Azide N_3^- Chlorate ion ClO_3^- Chlorite ion ClO_2^- Hypochlorite ClO^- Cyanide ion CN^- Hydroxide ion OH^- Nitrate ion NO_3^- Nitrite ion NO_2^- Perchlorate ion ClO_4^- Permanganate ion MnO_4^- Hydrogen carbonate ion HCO_3^- Dihydrogen phosphate ion H_2PO_4^- Thiocyanate SCN^- Bromate BrO_3^- Hypobromite BrO^- Iodate IO_3^- Periodate IO_4^- Hypoiodite IO^- Hydrogen sulfate HSO_4^- Hydrogen sulfite HSO_3^-
2-	Carbonate ion CO_3^{2-} Chromate ion CrO_4^{2-} Dichromate ion $\text{Cr}_2\text{O}_7^{2-}$ Hexafluorosilicate SiF_6^{2-} Hydrogen phosphate ion HPO_4^{2-} Peroxide ion O_2^{2-} Sulfate ion SO_4^{2-} Sulfite ion SO_3^{2-} Oxalate $\text{C}_2\text{O}_4^{2-}$ Selenate SeO_4^{2-} Silicate SiO_3^{2-} Tartrate $\text{C}_4\text{H}_4\text{O}_6^{2-}$ Tetraborate $\text{B}_4\text{O}_7^{2-}$ Thiosulfate $\text{S}_2\text{O}_3^{2-}$
3-	Phosphate ion PO_4^{3-} Phosphite PO_3^{3-} Arsenate ion AsO_4^{3-} Borate BO_3^{3-}
Positive charge	Ion name and formula
1+	Ammonium ion NH_4^+
2+	Mercury(I) ion Hg_2^{2+}

Common Acid Names

$\text{HC}_2\text{H}_3\text{O}_2$	acetic acid	HNO_3	nitric acid
CH_3COOH	acetic acid	H_3PO_4	phosphoric acid
H_2CO_3	carbonic acid	H_2SO_4	sulfuric acid
HCl	hydrochloric acid		