

NAME:

HONORS CHEMISTRY

SECTION:

Pairs/Check: Writing Names and Formulas

- A. You will work with a partner for this activity. The younger partner will do the even problems, and the older partner will do the odd problems. After completing 2 problems each, trade papers and check your partner's work. If a problem is correct, mark it with a check. Circle any incorrect answers. If you cannot agree on an answer, ask your teacher for assistance.
- B. Complete the following table. Consider the elements that make up the compound. A few are shown as examples. Be sure to include Roman numerals for Group II cations and parentheses as needed. (Remember, the group 1 metals, group 2 metals, aluminum, gallium, zinc and silver form only one ion!) Also, remember that prefixes are ONLY used for molecular substances (Type III binary compounds).

FORMULA	IONIC	MOLECULAR	NAME
Ru_2O_3	x		Ruthenium(III) oxide
SF_6		x	Sulfur hexafluoride
1.			Nickel(III) sulfite
2. N_2O_4			
3. Na_2O_2			
4.			Sulfur trioxide
5.			Silver phosphate
6. HgS			
7. $\text{Mn}(\text{CO}_3)_2$			
8.			Phosphorus trihydride
9. SeCl_2			
10. CuOH			
11.			Rhodium(III) acetate
12. $\text{Au}(\text{ClO}_2)_3$			
13.			Barium permanganate
14.			Diphosphorus pentoxide
15. RhCrO_4			
16. CBr_4			
17.			Lead(IV) cyanide
18.			Dibromine heptachloride
19. $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$			
20. N_2O_3			