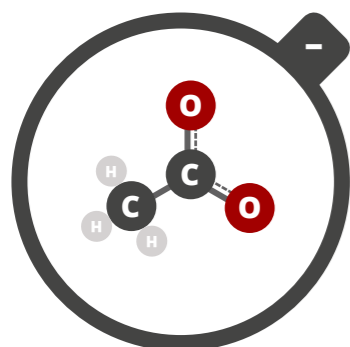


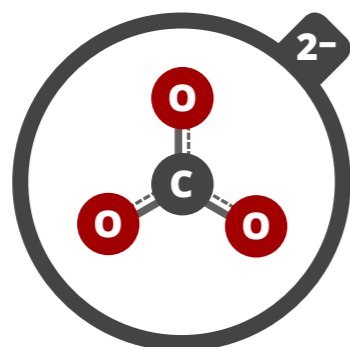
POLYATOMIC IONS: NAMES, FORMULAE & CHARGES

A polyatomic ion is a charged species consisting of two or more atoms covalently bonded together. Here's a guide to some of the most common examples!



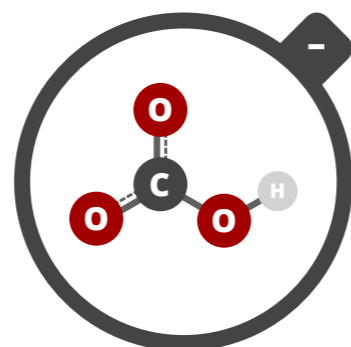
ACETATE

Formula: $C_2H_3O_2^-$



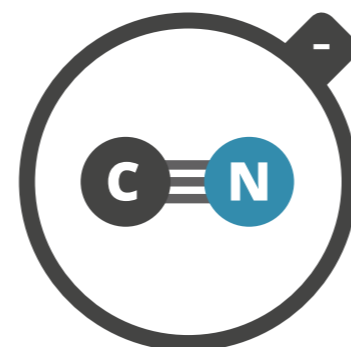
CARBONATE

Formula: CO_3^{2-}



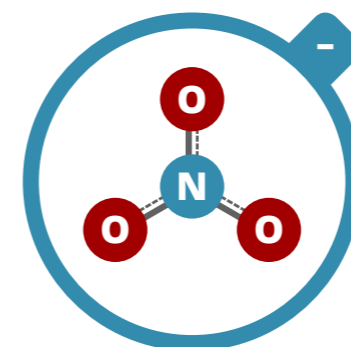
HYDROGEN CARBONATE

Formula: HCO_3^-



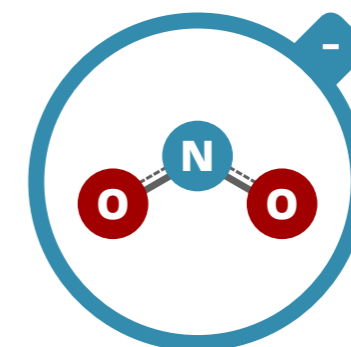
CYANIDE

Formula: CN^-



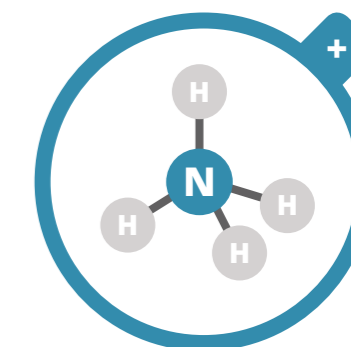
NITRATE

Formula: NO_3^-



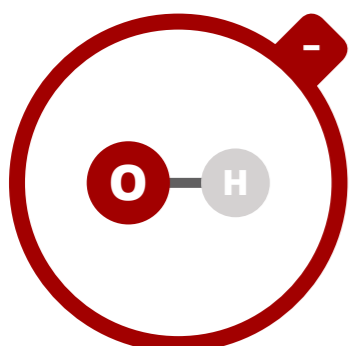
NITRITE

Formula: NO_2^-



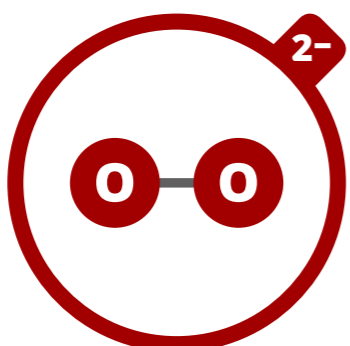
AMMONIUM

Formula: NH_4^+



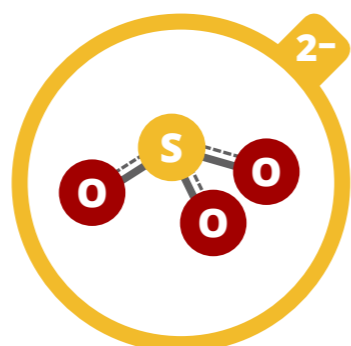
HYDROXIDE

Formula: OH^-



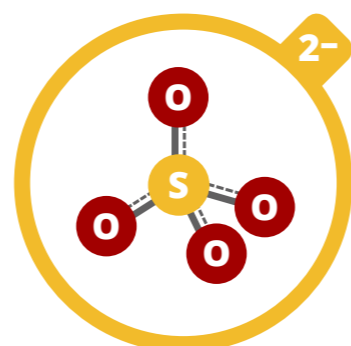
PEROXIDE

Formula: O_2^{2-}



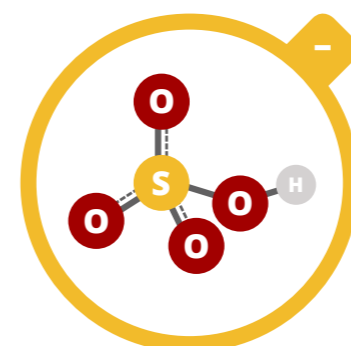
SULFITE

Formula: SO_3^{2-}



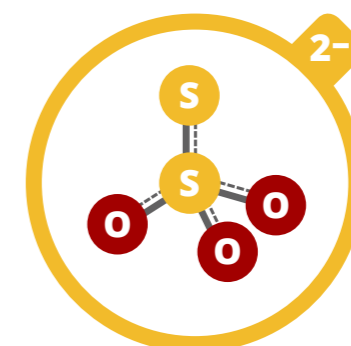
SULFATE

Formula: SO_4^{2-}



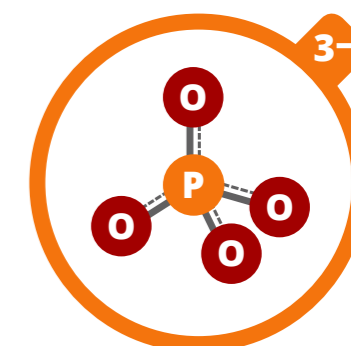
HYDROGEN SULFATE

Formula: HSO_4^-



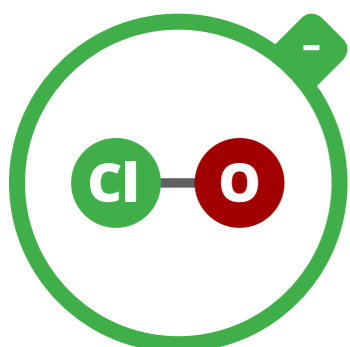
THIOSULFATE

Formula: $S_2O_3^{2-}$



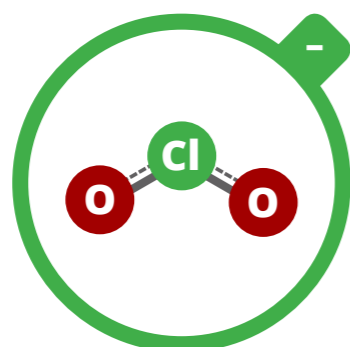
PHOSPHATE

Formula: PO_4^{3-}



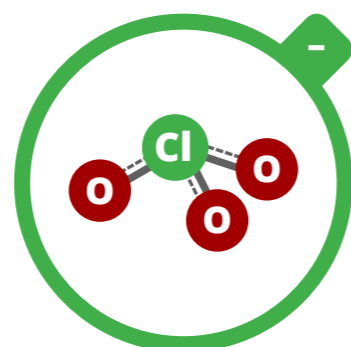
HYPOCHLORITE

Formula: ClO^-



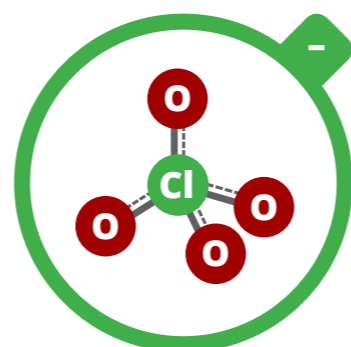
CHLORITE

Formula: ClO_2^-



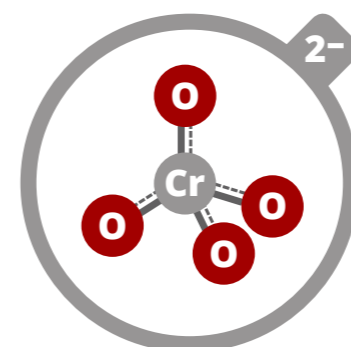
CHLORATE

Formula: ClO_3^-



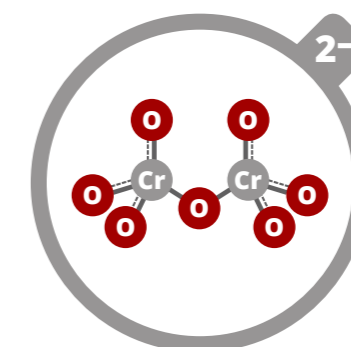
PERCHLORATE

Formula: ClO_4^-



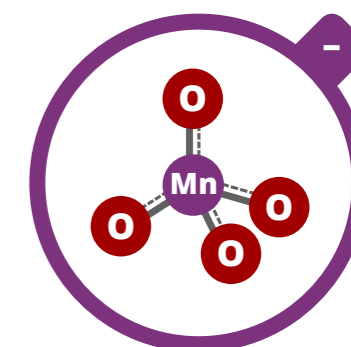
CHROMATE

Formula: CrO_4^{2-}



DICHROMATE

Formula: $Cr_2O_7^{2-}$



PERMANGANATE

Formula: MnO_4^-

