VSEPR Foldable

**Directions:**

Fold a piece of paper in half the long way.

Make four cuts to the fold, making 5 flaps.

Label the outside of each flap (1 label per flap):

Linear Trigonal planar

Bent Trigonal pyramidal

Tetrahedral

Cut out the boxes below.

On the back of each flap, attach the word description that correctly describes that molecular geometry. On the inside, attach the visual that corresponds to that molecular geometry. Then, attach the example molecule with that geometry.

When you are done, glue your foldable into your interactive notebook.

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| Two or three atoms that are connected in a straight line with a bond angle of 180o | Three atoms (but no lone pairs) are attached to a central atom, with 120o bond angles | Three atoms and one lone pair are attached to a central atom, with 107o bond angles | | Four atoms (but no lone pairs) are attached to a central atom, with 109.5o bond angles | Two atoms and two lone pairs are attached to a central atom, with a 104.5o bond angle |
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|  | | | Image result for methane structure | | |
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